

A. Subbiahpani,^a
D. Velmurugan,^{b*}
K. Ravikumar,^c N. Arumugam^d
and R. Raghunathan^d

^aDepartment of Physics, Presidency College (Autonomous), Chennai 600 005, India,

^bDepartment of Crystallography and Biophysics, University of Madras, Guindy Campus, Chennai 600 025, India, ^cLaboratory of X-ray Crystallography, Indian Institute of Chemical Technology, Hyderabad 500 007, India, and ^dDepartment of Organic Chemistry, University of Madras, Guindy campus, Chennai 600 025, India

Correspondence e-mail: d_velu@yahoo.com

Key indicators

Single-crystal X-ray study

$T = 293$ K

Mean $\sigma(\text{C}-\text{C}) = 0.003$ Å

R factor = 0.059

wR factor = 0.193

Data-to-parameter ratio = 17.2

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

4'-(4-Chlorophenyl)-2'-[1-(4-methoxyphenyl)-4-oxo-3-phenylazetidin-2-yl]-1'-methyl-2-phenyl-spiro[1,3-oxazole-4(5H),3'-pyrrolidin]-5-one

The pyrrolidine ring of the title molecule, $\text{C}_{35}\text{H}_{30}\text{ClN}_3\text{O}_4$, adopts a half-chair conformation; the oxazole and azetidine rings are essentially planar. Intermolecular $\text{C}-\text{H}\cdots\text{O}$ interactions are involved in the formation of dimers which are connected into a chain running along the [010] direction.

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Comment

The spiro-ring system is a frequently encountered structural motif in many pharmacologically relevant alkaloids. Pyrrolidine derivatives possess anti-influenza virus (Stylianakis *et al.*, 2003) and anticonvulsant (Obniska *et al.*, 2002) activity. Substituted pyrrolidine compounds have been found to have antimicrobial and antifungal activity against various pathogens (Amalraj *et al.*, 2003). Spiro-oxazole derivatives have been used in many natural product syntheses and have also been proved to be efficient precursors for many synthetic intermediates including γ -aminoalcohols, β -hydroxyketones, *etc.* (Kozikowski, 1984; Kanemasa & Tsuge, 1990). Spiro-oxazole compounds display interesting biological properties such as herbicidal, plant-growth regulatory and antitumour activities (Howe & Shelton 1990; De Amici *et al.*, 1990; Smietana *et al.*, 1999). The title compound, (I), was chosen for crystallographic study to establish its structure.

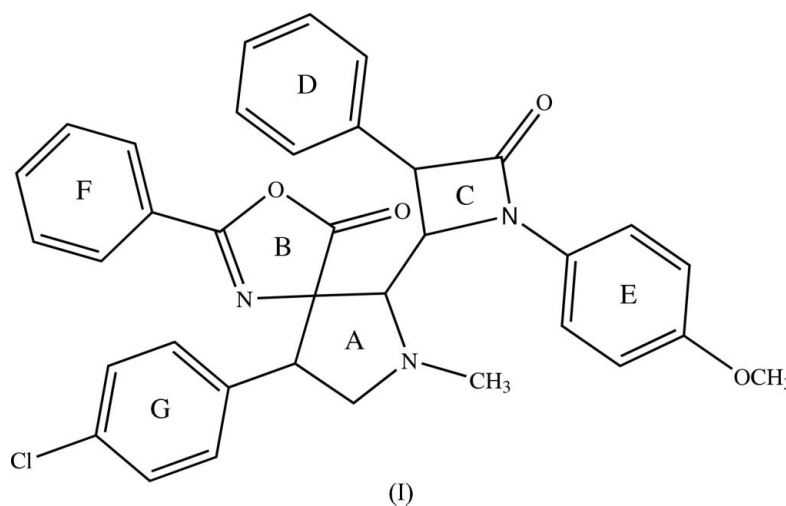


Fig. 1 shows a displacement ellipsoid plot of the title molecule, with the atomic numbering scheme. The pyrrolidine ring is in a half-chair conformation, with lowest displacement asymmetry parameter (Nardelli, 1983) $\Delta C_2(\text{C}1) = 0.009$ (1), and puckering parameters (Cremer & Pople, 1975) $q_2 = 0.425$ (2) Å and $\varphi = 123.2$ (2)°. The oxazole and azetidine rings are essentially planar, with maximum deviations of 0.043 Å for

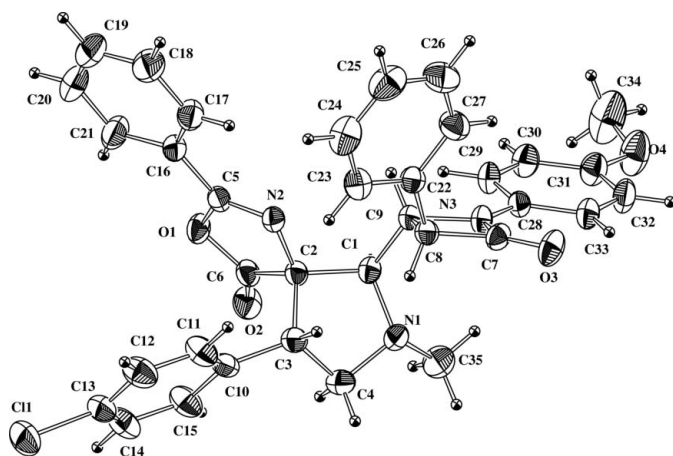


Figure 1

The molecular structure and atom-numbering scheme for (I). Displacement ellipsoids are drawn at the 30% probability level and H atoms are shown as small spheres of arbitrary radii.

C2 and 0.023 (2) Å for N3. The oxazole ring makes a dihedral angle of 2.1 (1)° with the attached phenyl ring.

The methoxy group is almost coplanar with the attached phenyl ring. The ketone atoms O2 and O3 deviate by 0.164 (2) and 0.046 (2) Å, respectively, from the least-squares planes of the oxazole and azetidine rings.

In the crystal structure, there is an intermolecular C32—H32···O4ⁱ hydrogen bond (symmetry code as in Table 2), resulting in the formation of dimers with graph-set motif $R_2^2(8)$ (Bernstein *et al.*, 1995). The dimers are linked by the C19—H19···O3ⁱⁱ hydrogen bond (Table 2), forming $C(13)$ chains running along the *b*-axis direction. A noteworthy feature of the crystal structure is the existence of a short intermolecular halogen—carbon contact, *i.e.* Cl1···C17ⁱⁱⁱ 3.226 (3) Å [symmetry code: (iii) $x - 1, y, z$].

Experimental

A mixture of *cis*-4-formyl-2-azetidinone (1 mmol), sarcosine (1 mmol) and 4-(*p*-chlorobenzylidene-2-phenyloxazol)-5-one (1 mmol) in dry toluene (15 ml) was refluxed for 12 h at 383 K, using a Dean–Stark apparatus. After completion of the reaction, as evidenced by thin-layer chromatography, the solvent was evaporated under reduced pressure. The crude product was purified by column chromatography using hexane–ethyl acetate (8.5:1.5) as eluent. The product was recrystallized from CHCl₃–hexane (1:2) by slow evaporation (m.p. 518 K).

Crystal data

C₃₅H₃₀ClN₃O₄
M_r = 592.07
 Triclinic, $P\bar{1}$
a = 10.0409 (10) Å
b = 13.2335 (13) Å
c = 13.4951 (13) Å
 α = 117.751 (2)°
 β = 94.976 (2)°
 γ = 98.865 (2)°

V = 1542.0 (3) Å³
Z = 2
D_x = 1.275 Mg m⁻³
 Mo *K*α radiation
 μ = 0.17 mm⁻¹
T = 293 (2) K
 Block, colourless
 0.38 × 0.25 × 0.16 mm

Data collection

Bruker SMART APEX CCD area-detector diffractometer
 ω scans
 Absorption correction: none
 9969 measured reflections

6680 independent reflections
 4825 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$
 $\theta_{\text{max}} = 28.0^\circ$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.059$
 $wR(F^2) = 0.193$
 $S = 0.91$
 6680 reflections
 388 parameters

H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.15P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.44 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\text{min}} = -0.29 \text{ e } \text{Å}^{-3}$

Table 1

Selected torsion angles (°).

C35—N1—C4—C3	−164.65 (18)	C34—O4—C31—C32	−171.2 (3)
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Table 2

Hydrogen-bond geometry (Å, °).

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
C32—H32···O4 ⁱ	0.93	2.58	3.476 (3)	161
C19—H19···O3 ⁱⁱ	0.93	2.59	3.295 (3)	133
C21—H21···O1	0.93	2.43	2.764 (3)	101
C33—H33···O3	0.93	2.51	3.114 (3)	123

Symmetry codes: (i) $-x + 3, -y, -z$; (ii) $x, y + 1, z$.

All H atoms were positioned geometrically and allowed to ride on their parent C atoms, with C—H = 0.93–0.98 Å and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{methyl C})$ or $1.2U_{\text{eq}}(\text{C})$.

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: ZORTEP (Zsolnai, 1997); software used to prepare material for publication: SHELXL97 and PARST (Nardelli, 1995).

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References

- Amalraj, A., Raghunathan, R., Sridevi Kumari, M. R. & Raman, N. (2003). *Bioorg. Med. Chem.* **11**, 407–419.
 Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
 Bruker (2001). SAINT (Version 6.28a) and SMART (Version 5.625). Bruker AXS Inc., Madison, Wisconsin, USA.
 Cremer, D. & Pople, J. A. (1975). *J. Am. Chem. Soc.* **97**, 1354–1358.
 De Amici, M., De Micheli, C. & Misani, V. (1990). *Tetrahedron*, **46**, 1975–1986.
 Howe, R. K. & Shelton, B. R. (1990). *J. Org. Chem.* **55**, 4603–4607.
 Kanemasa, S. & Tsuge, O. (1990). *Heterocycles*, **30**, 719–736.
 Kozikowski, A. P. (1984). *Acc. Chem. Res.* **17**, 410–416.
 Nardelli, M. (1983). *Acta Cryst.* **C39**, 1141–1142.
 Nardelli, M. (1995). *J. Appl. Cryst.* **28**, 659.
 Obniska, J., Zeic, A. & Zagorska, A. (2002). *Acta Pol. Pharm.* **59**, 209–213.

Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.

Smietana, M., Gouverneur, V. & Mioskowski, C. (1999). *Tetrahedron Lett.* **40**, 1291–1294.

Stylianakis, I., Kolocousis, A., Kolocousis, N. & Fytas, E. (2003). *Bioorg. Med. Chem. Lett.* **10**, 1699–1703.

Zsolnai, L. (1997). *ZORTEP*. University of Heidelberg, Germany.